Cambridge Assessment International Education

Cambridge International AS & A Level

CHEMISTRY **9701/33**

Paper 3 (Advanced Practical Skills 1) **May/June 2021**

MARK SCHEME

Maximum Mark: 40

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

Cambridge International is publishing the mark schemes for the May/June 2021 series for most Cambridge IGCSE™, Cambridge International A and AS Level components and some Cambridge O Level components.

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Generic Marking Principles

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptors for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently, e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

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GENERIC MARKING PRINCIPLE 5: Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).
GENERIC MARKING PRINCIPLE 6: Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

Science-Specific Marking Principles

1	Examiners should consider the context and scientific use of any keywords when awarding marks. Although keywords may be present, marks should not be awarded if the keywords are used incorrectly.
2	The examiner should not choose between contradictory statements given in the same question part, and credit should not be awarded for any correct statement that is contradicted within the same question part. Wrong science that is irrelevant to the question should be ignored.
3	Although spellings do not have to be correct, spellings of syllabus terms must allow for clear and unambiguous separation from other syllabus terms with which they may be confused (e.g. ethane / ethene, glucagon / glycogen, refraction / reflection).
4	The error carried forward (ecf) principle should be applied, where appropriate. If an incorrect answer is subsequently used in a scientifically correct way, the candidate should be awarded these subsequent marking points. Further guidance will be included in the mark scheme where necessary and any exceptions to this general principle will be noted.
5	‘List rule’ guidance For questions that require ***n*** responses (e.g. State **two** reasons …): • The response should be read as continuous prose, even when numbered answer spaces are provided. • Any response marked *ignore* in the mark scheme should not count towards ***n***. • Incorrect responses should not be awarded credit but will still count towards ***n***. • Read the entire response to check for any responses that contradict those that would otherwise be credited. Credit should **not** be awarded for any responses that are contradicted within the rest of the response. Where two responses contradict one another, this should be treated as a single incorrect response. • Non-contradictory responses after the first ***n*** responses may be ignored even if they include incorrect science.

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6 Calculation specific guidance

Correct answers to calculations should be given full credit even if there is no working or incorrect working, **unless** the question states 'show your working'.

For questions in which the number of significant figures required is not stated, credit should be awarded for correct answers when rounded by the examiner to the number of significant figures given in the mark scheme. This may not apply to measured values.

For answers given in standard form (e.g. $a \times 10^n$) in which the convention of restricting the value of the coefficient (a) to a value between 1 and 10 is not followed, credit may still be awarded if the answer can be converted to the answer given in the mark scheme.

Unless a separate mark is given for a unit, a missing or incorrect unit will normally mean that the final calculation mark is not awarded. Exceptions to this general principle will be noted in the mark scheme.

7 Guidance for chemical equations

Multiples / fractions of coefficients used in chemical equations are acceptable unless stated otherwise in the mark scheme.

State symbols given in an equation should be ignored unless asked for in the question or stated otherwise in the mark scheme.

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Question	Answer	Marks
1(a)	**I** The following data must be shown: • burette readings **and** titre for rough titration • 2 × 2 'box' showing both accurate burette readings	**1**
	II Headings and units correct for accurate titration table and headings match readings. • initial/start **and** (burette) reading / volume + unit • final/end **and** (burette) reading / volume + unit • titre **or** volume/**FA 5 and** used / added + unit Units: (cm^3) or / cm^3 or in cm^3 or cm^3 by every entry	**1**
	III All accurate burette readings to 0.05 cm^3	**1**

Question	Answer	Marks
1(a)	**IV** The **final** accurate titre recorded is within 0.10 cm^3 of any other accurate titre	**1**
	Accuracy (Q) marks Round burette readings to the nearest 0.05 cm^3. Check and correct titre subtractions where necessary. Examiner selects the best mean titre. Apply hierarchy: 2 identical, titres within 0.05 cm^3, titres within 0.10 cm^3 etc. Examiner subtracts (corrected) candidate's titre from Supervisor's titre. Write and ring supervisor's value next to the accurate titration table of each candidate, also candidate mean value (calculated by examiner) and difference between supervisor's and candidate's value, δ.	
	Award accuracy marks as follows: **V** Award if $\delta \leqslant 0.60\ cm^3$ **VI** Award if $\delta \leqslant 0.40\ cm^3$ **VII** Award if $\delta \leqslant 0.20\ cm^3$ (3)	**7**
1(b)	Correctly calculates mean accurate titre • Candidate must take the average of two (or more) titres that are within a total spread of not more than 0.20 cm^3. • Working/explanation must be shown **or** ticks must be put next to the two (or more) accurate readings selected. • The mean should be quoted to 2 dp and be rounded to nearest 0.01 cm^3.	**1**
1(c)(i)	Answers for **(c)(ii)** and **(iii)** to 3 or 4 sf	**1**
1(c)(ii)	Correctly calculates amount of I_2 = [**(b)** $\times$ 0.1] / [2 $\times$ 1000] (= **(b)** $\times$ 5 $\times$ 10^{-5} mol)	**1**

Question	Answer	Marks
1(c)(iii)	Correctly uses: If candidate's answer = 400 × **(c)(ii) and** working shown award both marks If candidate's answer = 400 × **(c)(ii)** no working shown award 1 mark If candidate's answer = 200 × **(c)(ii)** award 1 mark (I_2 / I^- confusion) If candidate's answer = 100 × **(c)(ii)** award 1 mark (250 / 25 instead of 1000 / 25) If candidate's answer = 40 × **(c)(ii)** award 1 mark (dilution factor missing)	**2**
1(d)(i)	moles of KI = 15 / 1000 × 0.5 = 7.5×10^{-3} mol moles I^- = **(c)(ii)** × 2 2 x **(c)(ii)** < 7.5×10^{-3} **OR** a comparison in words (**(c)(ii)** < 3.75×10^{-3})	**1**
1(d)(ii)	M1: Repeat the experiment using a greater volume of KI M2: If titre value is the same then KI was in excess **OR** M1: Add more KI at the endpoint M2: If mixture / it goes **black** / **blue-black** / **dark blue** then KI was not in excess **OR** M1: Add $AgNO_3$ **before** titrating M2: Yellow ppt observed shows presence of I^- / no yellow ppt shows I^- not in excess **OR** M1: Add more/extra NaC*l*O **when the titration mixture is yellow** M2: if the mixture goes darker yellow / brown etc., (this shows more iodine has been produced, as) KI is in excess. (ORA)	**2**

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Question	Answer	Marks
2(a)	**I** Table / list to include initial and final balance reading, initial and final thermometer reading, mass added and temperature change. **and** all entries with correct units: / °C, (g) or in °C or in degrees Celcius, in g or in gram(me)s **and** recorded in the space provided	**1**
	II All balance readings recorded to the same precision (minimum 1 dp) **and** all thermometer readings recorded to .0 or .5 °C.	**1**
	Accuracy marks Examiner checks and corrects mass of **FA 6** used and temperature fall (round thermometer readings to the nearest .0 or .5 ºC). Examiner calculates ΔT / mass to 2 dp for supervisor and candidate Write and ring supervisor's ratio next to the results table of each candidate, also candidate ratio and difference between supervisor's and candidate's value, δ.	
	Award accuracy marks as follows: **III** Award if $\delta \leqslant 0.40\ °C\ g^{-1}$ **IV** Award if $\delta \leqslant 0.20\ °C\ g^{-1}$ (2)	**4**
2(b)(i)	Correctly calculates heat energy change = 20 × 4.2 × temperature change **and** answer given to 2–4 sf	**1**
2(b)(ii)	M1: Correct display of mass of **FA 6** / moles of thiosulfate **or** **(b)(i)** / 47400 M2: Correct use of (correct mass **FA 6** × 47400) / ans **(b)(i)** **and** answer given to 2–4 sf	**2**

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Question	Answer	Marks
2(b)(iii)	Correct display x = [**(b)(ii)** – 158.**2**] / 18 **and** answer given to the nearest integer	**1**
2(c)	M1: temperature fall / change would be less **OR** heat energy change would be less M2: therefore **moles**/ **amount** of thiosulfate would be fewer / less **and** so M_r would be greater	**2**

Question	Answer	Marks
FA 4 is KI(aq); **FA 5** is NaS_2O_3(aq); **FA 7** is $MgCl_2 \cdot 6H_2O$; **FA 8** is $FeCl_3$ / H^+; **FA 9** is HCOOH(aq); **FA 10** is NaS_2O_3(aq); **FA 11** is HCl(aq)		
3(a)(i)	2 * = 1 mark (round down) **white** solid / powder / crystals (at start) * forms liquid/ melts / dissolves (in water of crystallization) * condensation / water droplets (on side of test-tube) / water vapour * **white** residue / solid / powder forms * **NOT** ppt (blue) litmus turns red *	**2**
3(a)(ii)	M1: Selects NaOH **and** NH_3 **ONLY** M2 & 3: + NaOH: white ppt * which is insoluble in **excess** * + NH_3: white ppt * which is insoluble in **excess** * 2 * = 1 mark (round down)	**3**
3(a)(iii)	+ $AgNO_3$: white ppt	**1**

Question	Answer	Marks
3(a)(iv)	cation: magnesium / Mg^{2+} **and** anion: chloride / Cl^- (from white ppt)	**1**
3(b)	**Test 1**. M1: + **FA 4** yellow (solution) turns darker yellow / yellow-brown / orange-brown / red-brown / brown M2: + starch – (turns) black / blue-black / dark blue **Test 2**. M3: + **FA 5** solution turns purple **and** colour fades on standing	**3**
3(c)(i)	**Test 1** **FA 9** effervescence / bubbling / fizzing * **FA 10** no reaction / no effervescence / bubbling / fizzing * **FA 11** faster effervescence / bubbling / fizzing * gas / H_2 pops with a lighted splint with **FA 9** and / or **FA 11** * **Test 2** **FA 9** $KMnO_4$ decolourised **OR** purple to / turns colourless * **FA 10** $KMnO_4$ decolourised **OR** purple to / turns colourless * **FA11** no reaction **OR** solution remains purple * 2* = 1 mark (round down)	**3**

Question	Answer	Marks
3(c)(ii)	**FA 9** = HCOOH / methanoic acid; **FA 10** = Na_2SO_3 / sodium sulfite / sodium sulfate(IV) **FA 11** = HCl / hydrochloric acid 1 mark for one correct 2 marks for all three correct	**2**